

Redox

Chemistry

Oxidation and reduction, and the rules for oxidation numbers.

Rules for assigning oxidation numbers.

Rule	Value
Element in its standard state	0
Monatomic ion	Its charge
Oxygen, usually	-2
Oxygen in peroxides	-1
Hydrogen with non-metals	+1
Hydrogen with metals	-1
Group 1 / Group 2	+1 / +2
Sum in a neutral compound	0

The oxidising agent is itself reduced, and vice versa. That reversal is the commonest exam trap.