

Period 3

Chemistry

Trends across a single period, in properties and in oxide behaviour.

Across period 3, sodium to argon.

Property	Trend
Atomic radius	Decreases
Ionisation energy	Increases, with dips at Al and S
Melting point	Peaks at silicon
Electronegativity	Increases
Oxide character	Basic → amphoteric → acidic
Bonding in oxides	Ionic → covalent

Silicon's melting point peak is structural: it is a giant covalent lattice, while the others are metallic or simple molecular.