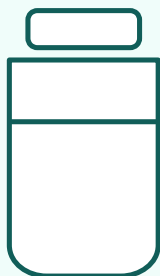


Ionic Bonding

10th Grade

Electrons handed over, not shared — and the properties that follow.



In an ionic bond a metal gives electrons to a non-metal. Both end up with full outer shells and opposite charges, and the attraction holds them together.

BETWEEN
metal + non-metal

ELECTRONS
transferred

- Metals lose electrons and become positive ions.
- Non-metals gain electrons and become negative ions.
- Group 1 forms +1; group 2 forms +2; group 17 forms -1.
- Ionic compounds form giant lattices, not molecules.
- High melting points, and they conduct when molten or dissolved.

LOOK FOR IT

Na loses one electron to Cl. Na^+ and Cl^- attract, forming NaCl.

Solid ionic compounds do not conduct — the ions are locked in place and cannot move.