

Enthalpy

Chemistry

Exothermic and endothermic, and the sign convention.

Standard enthalpy change definitions.

Term	Is
ΔH formation	1 mol of compound from its elements
ΔH combustion	1 mol burnt completely in oxygen
ΔH neutralisation	1 mol of water from acid and base
ΔH atomisation	1 mol of gaseous atoms
Bond enthalpy	Breaking 1 mol of bonds, gaseous
Lattice enthalpy	1 mol of ionic solid from gaseous ions

Breaking bonds is endothermic; making them is exothermic. $\Delta H = \text{bonds broken} - \text{bonds made}$.