

Concentration

Chemistry

Converting between the units concentration appears in.

Conversions between common units.

From → to	Do
$\text{mol dm}^{-3} \rightarrow \text{g dm}^{-3}$	$\times$ molar mass
$\text{g dm}^{-3} \rightarrow \text{mol dm}^{-3}$	$\div$ molar mass
$\text{cm}^3 \rightarrow \text{dm}^3$	$\div 1000$
$\text{dm}^3 \rightarrow \text{m}^3$	$\div 1000$
$\text{ppm} \rightarrow \text{mg dm}^{-3}$	Equal, for dilute aqueous
$\% \text{ w/v} \rightarrow \text{g dm}^{-3}$	$\times 10$

1 mol dm⁻³ = 1 M = 1 molar. All three notations appear and mean the same thing.