

Acids & Bases

Chemistry

The three definitions, and what each adds.

Increasing generality.

Theory	Acid is	Base is
Arrhenius	Produces H^+ in water	Produces OH^- in water
Brønsted-Lowry	Proton donor	Proton acceptor
Lewis	Electron pair acceptor	Electron pair donor

Brønsted-Lowry introduces conjugate pairs: an acid and its conjugate base differ by exactly one proton.